



Teaching Guide				
Identifying Data				2016/17
Subject (*)	Química	Code	770G02004	
Study programme	Grao en Enxeñaría Eléctrica			
Descriptors				
Cycle	Period	Year	Type	Credits
Graduate	1st four-month period	First	FB	6
Language	Spanish			
Teaching method	Face-to-face			
Prerequisites				
Department	Química Analítica			
Coordinador	Gonzalez Rodriguez, Maria Victoria	E-mail	victoria.gonzalez.rodriguez@udc.es	
Lecturers	Alonso Rodriguez, Elia Gonzalez Rodriguez, Maria Victoria	E-mail	elia.alonso@udc.es victoria.gonzalez.rodriguez@udc.es	
Web				
General description	Introdución aos fundamentos científicos da química en relación coas súas aplicacións tecnolóxicas			

Study programme competences / results	
Code	Study programme competences / results

Learning outcomes			
Learning outcomes	Study programme competences / results		
Utilize the basic principles of general chemistry, organic chemistry and inorganic chemistry. Apply the basic laws governing reactions: thermodynamics, kinetics and equilibrium.	A8		
Solve problems and analyze results.		B1	
Work and learn independently and with initiative.		B2 B4 B6	
Using information resources , applying ICT and transmit knowledge in the exercise of the profession.		B2 B6	C3

Contents	
Topic	Sub-topic
Topic 1. Basics of Chemistry.	<ul style="list-style-type: none"> - Stoichiometry. Theoretical and Percentage Yields. Limiting Reactant. - Atoms. The Quantum Mechanical Model. - Periodic Table of the Elements. - Chemical Bond. Main types of chemical bonds: ionic, covalent, metallic. Intermolecular Forces.
Topic 2. Thermochemistry	<ul style="list-style-type: none"> - Heats of Chemistry Reaction - Enthalpy - Calorimetry - Introduction to thermodynamics
Topic 3. Rates of Reaction	<ul style="list-style-type: none"> - Reaction Rates - Reaction Rates Equation - Dependence of Rate on Concentration - Activation energy - Catalysis - Mechanism



Topic 4. Chemical Equilibrium	<ul style="list-style-type: none"> - Chemical Equilibrium. The Equilibrium Constant. - Gaseous Reactions. Le Chatelier's Principle - Acid-Base Equilibria
Topic 5. Electrochemistry I	<ul style="list-style-type: none"> - Oxidation -Reduction Reactions. Balancing - Standard Electrode Potentials - Spontaneity from Electrode Potentials - Nernst Equation
Topic 6. Electrochemistry II	<ul style="list-style-type: none"> - Voltaic Cells. Batteries - Electrolysis. Stoichiometry of Electrolysis
Topic 7. Corrosion	<ul style="list-style-type: none"> - Concept - Corrosion process and influence factors - Methods to protect metals from corrosion - Atmospheric Corrosión - Marine Corrosion
Topic 8. Organic Chemistry	<ul style="list-style-type: none"> - Introduction to Organic Chemistry - Functional Groups - Nomenclature - Isomers - Main types of organic reactions
Topic 9. Organic Chemistry Applied to Engineering	<ul style="list-style-type: none"> - Carbon - Oil - Gas - Biomass - Polymers
Topic 10. Introduction to Industrial Chemistry	<ul style="list-style-type: none"> - Engineering Process - Mass Balance
Topic 11. Inorganic Chemistry Applied to Engineering	<ul style="list-style-type: none"> - Metallurgy - Industrial Inorganic Compounds: Synthesis - Main Technologic Inorganic Materials: Semiconductors, Optic Fiber, Ceramic, Superconductors
Topic 12. Introduction to Instrumental Techniques for Industrial Analysis	<ul style="list-style-type: none"> - Classification of Instrumental Techniques - Quality Parameters in the Analytical Laboratory - Calibration - Significant Digits

Planning				
Methodologies / tests	Competencies / Results	Teaching hours (in-person & virtual)	Student's personal work hours	Total hours
Guest lecture / keynote speech	A8	21	29.4	50.4
Problem solving	A8 B1	20	38	58
Multiple-choice questions	A8 B4 B6 C3	1	4	5
Laboratory practice	A8 B6	5	5	10
Supervised projects	B2 B4 B6 C3	3	6	9
Objective test	A8 B1	4	12	16
Personalized attention		1.6	0	1.6

(*)The information in the planning table is for guidance only and does not take into account the heterogeneity of the students.

Methodologies	
Methodologies	Description



Guest lecture / keynote speech	Participants take notes and make questions
Problem solving	Participants apply rules, write mathematical relationships and analyze results
Multiple-choice questions	Participants complete online exercises and individual self-assessment is available
Laboratory practice	Participants perform an experiment following a written procedure and write a report
Supervised projects	Participants summarize and discuss information
Objective test	Participants answer questions and problems

Personalized attention

Methodologies	Description
Supervised projects	<p>Reviewing the development of intermediate and final stages of supervised projects</p> <p>Resolving specific issues</p> <p>Students being recognized officially as partial-time and entitled not to attend the lectures will be attended in a tutorships regime (set hour with teacher in advance).</p>

Assessment

Methodologies	Competencies / Results	Description	Qualification
Problem solving	A8 B1	Resolution of exercises and ability to explain them in the classroom	10
Laboratory practice	A8 B6	Carry out the laboratory practices and reports and ability to work collaboratively	5
Supervised projects	B2 B4 B6 C3	Elaboration of supervised projects and presentation in the classroom. Performing an activity and objective test.	10
Objective test	A8 B1	<p>A first test (theory and problems) will be carried out about half of the semester. The subject taught until then will be evaluated. At the end of course, a partial second test (theory and problems) will be performed for students who have passed the first test. Simultaneously a global test (theory and problems) will be performed for students who have not approved the first test.</p> <p>Each test consists of two independent parts, being necessary to obtain a minimum score on each part to compensate:</p> <ul style="list-style-type: none"> - Theory, maximum score 4 points, minimum score 1.5 points to compensate. - Problems, maximum score 3 points, 1 point minimum to compensate score. 	70
Multiple-choice questions	A8 B4 B6 C3	Solving self-assessment online exercises proposed before test	5

Assessment comments

A minimum of 75% of the laboratory practical classes have to be carried out by each student to be evaluated.

A minimum mark of 3 points is requested in the test to take into account the other marks.

For students being recognized officially as partial-time and entitled not to attend the lectures, the final exam represent 80% of the final grade and supervised projects 20%.

Sources of information



Basic	<ul style="list-style-type: none">- CHANG (2002). Química . Interamericana. Mc Graw - Hill. 7ª Edición- http://eup.cdf.udc.es (). .- McMurry, Fay (2009). Química General . Prentice Hall- PÉREZ IGLESIAS, J. y SECO LAGO, H.M. (2006). Experimentos de química. Aplicaciones a la vida cotidiana . Badajoz. Editorial Filarias- VINAGRE F., VAZQUEZ DE MIGUEL L.M. (1996). Fundamentos y problemas de química . Alianza, 4ª Ed.- Petrucci, Ralph H. (2011). Química general: principios y aplicaciones modernas. Prentice Hall
Complementary	<ul style="list-style-type: none">- WILLIS (1995). Resolución de Problemas de Química General . Reverté- José Vale Parapar y col. (2004). Problemas resueltos de Química para Ingeniería . Thomson- KOTZ, TREICHEL, HARMAN (2003). Química y reactividad química . Thomson Ed. 5º Ed.- PAZ, M.; CASTRO, F. y MIRO, J. (1995). Química . Madrid.Ed.UNED- PETERSON (2012). Fundamentos de nomenclatura química . Reverte- Skoog, Douglas A (2007). Principios de análisis instrumental . Santa Fe : Cengage Learning

Recommendations

Subjects that it is recommended to have taken before

Subjects that are recommended to be taken simultaneously

Subjects that continue the syllabus

Enxeñaría Medioambiental/770G01014

Other comments

(*)The teaching guide is the document in which the URV publishes the information about all its courses. It is a public document and cannot be modified. Only in exceptional cases can it be revised by the competent agent or duly revised so that it is in line with current legislation.