



| Teaching Guide | | | | |
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| Identifying Data | | | 2022/23 | |
| Subject (*) | Química | Code | 770G02004 | |
| Study programme | Grao en Enxeñaría Eléctrica | | | |
| Descriptors | | | | |
| Cycle | Period | Year | Type | Credits |
| Graduate | 1st four-month period | First | Basic training | 6 |
| Language | Spanish | | | |
| Teaching method | Face-to-face | | | |
| Prerequisites | | | | |
| Department | Química | | | |
| Coordinador | González Rodríguez, María Victoria | E-mail | victoria.gonzalez.rodriguez@udc.es | |
| Lecturers | Alonso Rodríguez, Elia González Rodríguez, María Victoria | E-mail | elia.alonso@udc.es victoria.gonzalez.rodriguez@udc.es | |
| Web | | | | |
| General description | Introduction to the scientific foundations of chemistry in relation to their technological applications | | | |

| Study programme competences | |
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| Code | Study programme competences |
| A8 | Capacidade para comprender e aplicar os principios e coñecementos básicos da química xeral, química orgánica e inorgánica e as súas aplicacións na enxeñaría. |
| B1 | Capacidade de resolver problemas con iniciativa, toma de decisións, creatividade e razoamento crítico. |
| B2 | Capacidade de comunicar e transmitir coñecementos, habilidades e destrezas no campo da enxeñaría industrial. |
| B4 | Capacidade de traballar e aprender de forma autónoma e con iniciativa. |
| B6 | Capacidade de usar adecuadamente os recursos de información e aplicar as tecnoloxías da información e as comunicacións na enxeñaría. |
| B7 | Capacidade para traballar de forma colaborativa e de motivar un grupo de traballo. |
| C3 | Utilizar as ferramentas básicas das tecnoloxías da información e as comunicacións (TIC) necesarias para o exercicio da súa profesión e para a aprendizaxe ao longo da súa vida. |

| Learning outcomes | | | |
|--|----|----------|-----------------------------|
| Learning outcomes | | | Study programme competences |
| Utilize the basic principles of general chemistry, organic chemistry and inorganic chemistry. | A8 | B7 | C3 |
| Apply the basic laws governing reactions: thermodynamics, kinetics and equilibrium. | A8 | | C3 |
| Solve problems and analyze results. | A8 | B7 | C3 |
| Adequately apply theoretical concepts in the laboratory through the correct and safe use of basic material and equipment | | B1 B4 | |
| Use rigorous language in chemistry | | B2 | |
| Present and interpret data and results | | B6 B7 | |

| Contents | |
|-------------------------------|--|
| Topic | Sub-topic |
| Unity 1. Chemistry basics | Includes topic 1 |
| Topic 1. Basics of Chemistry. | - Stoichiometry. Theoretical and Percentage Yields. Limiting Reactant. - Atoms. The Quantum Mechanical Model. - Periodic Table of the Elements. - Chemical Bond. Main types of chemical bonds: ionic, covalent, metallic. Intermolecular Forces. |



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| Unity 2. Thermochemistry | Includes topic 2 |
| Topic 2. Thermochemistry | <ul style="list-style-type: none">- Heats of Chemistry Reaction- Enthalpy- Calorimetry- Introduction to thermodynamics |
| Unity 3. Rates of Reaction | Includes topic 3 |
| Topic 3. Rates of Reaction | <ul style="list-style-type: none">- Reaction Rates- Reaction Rates Equation- Dependence of Rate on Concentration- Activation energy- Catalysis- Mechanism |
| Unity 4. Chemical Equilibrium | Includes topic 4 |
| Topic 4. Chemical Equilibrium | <ul style="list-style-type: none">- Chemical Equilibrium. The Equilibrium Constant.- Gaseous Reactions. Le Chatelier's Principle- Acid-Base Equilibria |
| Unity 5. Electrochemistry | Includes topics 5, 6 and 7 |
| Topic 5. Electrochemistry I | <ul style="list-style-type: none">- Oxidation -Reduction Reactions. Balancing- Standard Electrode Potentials- Spontaneity from Electrode Potentials- Nernst Equation |
| Topic 6. Electrochemistry II | <ul style="list-style-type: none">- Voltaic Cells. Batteries- Electrolysis. Stoichiometry of Electrolysis |
| Topic 7. Corrosion | <ul style="list-style-type: none">- Concept- Corrosion process and influence factors- Methods to protect metals from corrosion- Atmospheric Corrosión- Marine Corrosion |
| Unity 6. Principles of Organic Chemistry | Includes topic 8 |
| Topic 8. Organic Chemistrya | <ul style="list-style-type: none">- Introduction to Organic Chemistry- Functional Groups- Nomenclature- Isomers- Main types of organic reactions |
| Unity 7. Organic and Inorganic Chemistry Applied to Engineering | Includes topics 9 and 10 |
| Topic 9. Organic Chemistry Applied to Engineering | <ul style="list-style-type: none">- Carbon- Oil- Gas- Biomass- Polymers |
| Topic 10. Inorganic Chemistry Applied to Engineering | <ul style="list-style-type: none">- Metallurgy- Industrial Inorganic Compounds: Synthesis- Main Technologic Inorganic Materials: Semiconductors, Optic Fiber, Ceramic, Superconductors |
| Unity 8. Bases of Industrial Chemistry: Mass Balance | Includes topic 8 |
| Topic 11. Introduction to Industrial Chemistry | <ul style="list-style-type: none">- Engineering Process- Mass Balance |
| Unity 9. Principles of Instrumental Analysis | Includes topic 12 |



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| Topic 12. Introduction to Instrumental Techniques for Industrial Analysis | <ul style="list-style-type: none"> - Classification of Instrumental Techniques - Quality Parameters in the Analytical Laboratory - Calibration - Significant Digits |
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| Planning | | | | |
|--------------------------------|----------------|----------------------|-------------------------------|-------------|
| Methodologies / tests | Competencies | Ordinary class hours | Student's personal work hours | Total hours |
| Guest lecture / keynote speech | A8 | 30 | 43.5 | 73.5 |
| Problem solving | B1 B7 | 20 | 28.4 | 48.4 |
| Laboratory practice | A8 B4 B6 B7 C3 | 5 | 2.5 | 7.5 |
| Supervised projects | B2 B7 C3 | 1 | 2 | 3 |
| Objective test | A8 B1 | 4 | 12 | 16 |
| Personalized attention | | 1.6 | 0 | 1.6 |

(*)The information in the planning table is for guidance only and does not take into account the heterogeneity of the students.

| Methodologies | |
|--------------------------------|---|
| Methodologies | Description |
| Guest lecture / keynote speech | Participants take notes and make questions |
| Problem solving | Participants apply rules, write mathematical relationships and analyze results |
| Laboratory practice | Participants perform an experiment following a written procedure and write a report |
| Supervised projects | Participants summarize and discuss information |
| Objective test | Participants answer questions and problems |

| Personalized attention | |
|------------------------|---|
| Methodologies | Description |
| Supervised projects | <p>Reviewing the development of intermediate and final stages of supervised projects</p> <p>Resolving specific issues</p> <p>Students being recognized officially as partial-time and entitled not to attend the lectures will be attended in a tutorships regime (set hour with teacher in advance).</p> |

| Assessment | | | |
|---------------------|----------------|---|---------------|
| Methodologies | Competencies | Description | Qualification |
| Problem solving | B1 B7 | Resolution of questions, exercises and ability to explain them in the classroom | 20 |
| Laboratory practice | A8 B4 B6 B7 C3 | Carry out the laboratory practices and reports and ability to work collaboratively | 10 |
| Supervised projects | B2 B7 C3 | Elaboration of supervised projects and presentation in the classroom. Performing an activity and objective test. | 10 |



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| Objective test | A8 B1 | <p>A first test (theory and problems) will be carried out about half of the semester. The subject taught until then will be evaluated. At the end of course, a partial second test (theory and problems) will be performed for students who have passed the first test. Simultaneously a global test (theory and problems) will be performed for students who have not approved the first test.</p> <p>Each test consists of two independent parts, being necessary to obtain a minimum score on each part to compensate:</p> <ul style="list-style-type: none"> - Theory, maximum score 3 points, minimum score 1.25 points to compensate. - Problems, maximum score 3 points, 1.25 points minimum to compensate score. | 60 |
|----------------|-------|--|----|

Assessment comments

A minimum of 75% of the laboratory practical classes have to be carried out by each student to be evaluated.

A minimum mark of 2.5 points is requested in the test to take into account the other marks.

For the evaluation of the second opportunity, these same continuous evaluation activities can be carried out as during the course except for laboratory practices and instead, some questionnaires can be carried out in Moodle

For students being recognized officially as partial-time and entitled not to attend the lectures, the final exam represents 80% of the final grade and supervised projects 20%.

The fraudulent performance of the tests or evaluation activities will directly imply the failure grade '0' in the matter in the corresponding call, thus invalidating any qualification obtained in all the evaluation activities for the extraordinary call

For 2010 Plan students, who explicitly renounce continuous assessment will be evaluated by the grade obtained in the final exam (100%)

Sources of information

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|----------------------|---|
| Basic | <ul style="list-style-type: none"> - CHANG (2002). Química. Interamericana. Mc Graw - Hill. 7ª Edición - http://eup.cdf.udc.es (). - McMurry, Fay (2009). Química General. Prentice Hall - PÉREZ IGLESIAS, J. y SECO LAGO, H.M. (2006). Experimentos de química. Aplicaciones a la vida cotidiana. Badajoz. Editorial Filarias - VINAGRE F., VAZQUEZ DE MIGUEL L.M. (1996). Fundamentos y problemas de química. Alianza, 4ª Ed. - Petrucci, Ralph H. (2011). Química general: principios y aplicaciones modernas. Prentice Hall |
| Complementary | <ul style="list-style-type: none"> - WILLIS (1995). Resolución de Problemas de Química General. Reverté - José Vale Parapar y col. (2004). Problemas resueltos de Química para Ingeniería. Thomson - KOTZ, TREICHEL, HARMAN (2003). Química y reactividad química. Thomson Ed. 5º Ed. - PAZ, M.; CASTRO, F. y MIRO, J. (1995). Química. Madrid. Ed. UNED - PETERSON (2012). Fundamentos de nomenclatura química. Reverte - Skoog, Douglas A (2007). Principios de análisis instrumental. Santa Fe: Cengage Learning |

Recommendations

Subjects that it is recommended to have taken before

Subjects that are recommended to be taken simultaneously

Subjects that continue the syllabus

Environmental Engineering/770G01014

Other comments



Recommendations Sustainability Environment, Person and Gender Equality:1.

The delivery of the works (supervised work) that

are carried out in this matter will be done in the following way:

- 1.1. It will be delivered in virtual format and / or computer support
- 1.2. In the case of having to print something on paper, it will be made on recycled and double-sided paper. Drafts will not be printed, only the final version.
2. It must make a sustainable use of resources and the prevention of negative impacts on the natural environment. It will be encouraged that the materials that are discarded in the matter (papers, plastics) are thrown in the respective containers enabled in the streets for such purpose.
3. It will try to convey to students the importance of ethical principles related to the values ??of sustainability so that they apply not only in the classroom, but in personal and professional behaviors.
4. The gender perspective must be incorporated in this subject, so the works delivered by the students and the material prepared by the teacher must use non-sexist language.
5. It will facilitate the full integration of students who for physical, sensory, psychic or sociocultural reasons, experience difficulties to an adequate, equal and profitable access to university life.

(*)The teaching guide is the document in which the URV publishes the information about all its courses. It is a public document and cannot be modified. Only in exceptional cases can it be revised by the competent agent or duly revised so that it is in line with current legislation.