

		Teachin	g Guide		
Identifying Data				2017/18	
Subject (*)	Physical Chemistry 1 Code			610G01016	
Study programme	Grao en Química			1	
	-	Desc	riptors		
Cycle	Period	Ye	ear	Туре	Credits
Graduate	1st four-month period	Sec	cond	Obligatoria	6
Language	SpanishEnglish				
Teaching method	Face-to-face				
Prerequisites					
Department	Química				
Coordinador	Rodriguez Barro, Pilar		E-mail	pilar.rbarro@udo	c.es
Lecturers	Rodriguez Barro, Pilar		E-mail	pilar.rbarro@udc.es	
	Vilariño Barreiro, Maria Teresa teresa.vilarino@udc.es			udc.es	
Web					
General description	al description This course deals with the microscopic behavior of matter. It presents an introduction to quantum mechanics and statistical				
	thermodynamics. The course begin	ns with an exa	amination of the key	concepts, the basic p	principles and the formulation of
	quantum theory and applications to simple systems ? the particle in a box, the harmonic oscillator, the rigid rotor and the				
	hydrogen atom. It continues with a discussion of atomic structure and atomic spectra. The final lectures on quantum				
	chemistry cover applications to chemical bonding including valence bond and molecular orbital theory and molecular				
	structure.				
	The last part of the course covers elementary statistical mechanics that allows one to study the methodology to calculate				
	macroscopic properties of equilibrium systems from molecular properties.				
	(English lecturer: Teresa Vilariño)				

	Study programme competences
Code	Study programme competences
A1	Ability to use chemistry terminology, nomenclature, conventions and units
A8	Knowledge of principles of quantum mechanics and atomic and molecular structure
A12	Ability to relate macroscopic properties of matter to its microscopic structure
A14	Ability to demonstrate knowledge and understanding of concepts, principles and theories in chemistry
A15	Ability to recognise and analyse new problems and develop solution strategies
A16	Ability to source, assess and apply technical bibliographical information and data relating to chemistry
A21	Understanding of qualitative and quantitative aspects of chemical problems
B2	Effective problem solving
B3	Application of logical, critical, creative thinking
B5	Teamwork and collaboration
C1	Ability to express oneself accurately in the official languages of Galicia (oral and in written)
C3	Ability to use basic information and communications technology (ICT) tools for professional purposes and learning throughout life

Learning outcomes			
Learning outcomes		Study programme	
		competences	
To know the principles of quantum chemistry.	A1	B2	C3
	A8	B5	
	A14		
	A15		
	A16		



To know the principles of statistical thermodynamics.	A1	B2	C3
	A12	B5	
	A14		
	A16		
To be able to perform calculations independently, even when a computer is needed.	A1	B2	C1
	A14	B5	C3
	A15		
	A16		
	A21		
To acquire literature search skills to be able to search for and use scientific literature.	A14	B3	C1
	A15		C3
	A16		
	A21		
To acquire skills in the use of computer tools to solve problems.	A8	B2	C3
	A15	B3	

Contents				
Торіс	Sub-topic			
QUANTUM CHEMISTRY				
1. Postulates of quantum mechanics	- Postulate 1: the state of a quantum-mechanical system is completely specified by its			
	wave function.			
	- Postulate 2: quantum-mechanical operators represent classical mechanical			
	variables.			
	- Postulate 3: eigenvalue equation.			
	- Postulate 4: average value.			
	- Postulate 5: time-dependent Schrödinger equation.			
2. Translational motion: a particle in a box.	- A particle in a one-dimensional box: wave functions and energy levels.			
	- A particle in two and more dimensions:separation of variables and degeneracy.			
3. Vibrational motion: the harmonic oscillator.	- Quantum mechanical model: wave functions and energy levels.			
	- The harmonic oscillator as a model for a vibrating diatomic molecule.			
	- Anharmonicity.			
4. Rotational motion: rigid rotator.	- Motion of a particle in a ring.			
	- Wave functions. Spherical harmonics.			
	- Rotational energy: energy levels.			
	- The quantization of angular momentum.			
5. Hydrogenic atoms.	- Formulation of the Schrödinger equation.			
	- Atomic orbitals and their energies.			
	- The radial probability distribution function.			
	- The lineal combination of degenerate wavefunctions.			
	- Zeeman effect.			
6. Aproximation methods.	- Perturbation theory.			
	- Variational method.			
	- Lineal variational trial functions: secular determinant.			
7. Many-electron atoms.	- Helium atom.			
	- Spin angular moment.			
	- Pauli exclusion principle.			
	- Periodic Table.			
8. Atomic spectroscopy.	- Electron configuration of atoms.			
	- Total orbital angular moment: Russell-Saunders coupling and jj coupling.			
	- Term symbols. Hund's rules. Selection rules.			



9. The chemical bond: the hydrogen molecule-ion.	- The Born-Oppenheimer approximation.
	- Molecular orbital theory and valence-bond theory.
	- Molecular orbital treatment of hydrogen molecule-ion.
10. Diatomic molecules.	- General considerations for bond formation.
	- Homonuclear diatomic molecules.
	- Heteronuclear diatomic molecules. Polar bonds and electronegativity.
11. Conjugated and aromatic molecules.	- Semi-empirical methods.
	- Pi-electron approximation.
	- Free electron molecular orbital theory.
	- The Hückel approximation.
STATISTICAL THERMODYNAMICS	
12. Foundations of statistical thermodynamics.	- Fundamentals of statistical mechanics.
	- Basis of statistical thermodynamics.
	- Statistical thermodynamics of ideal gases.
	- Statistical interpretation of the thermodynamic properties of solids.

Planning					
Methodologies / tests	Competencies	Competencies Ordinary class		Total hours	
		hours	work hours		
Guest lecture / keynote speech	A1 A8 A12 A21	28	56	84	
Seminar	A14 A15 B2 B3	10	25	35	
Laboratory practice	A1 A8 A21 C1 C3	10	5	15	
Supervised projects	A1 A8 A16 B2 B3 B5	0	10	10	
	C1 C3				
Objective test	A1 A8 A14	1	0	1	
Objective test	A1 A8 A14	1	0	1	
Mixed objective/subjective test	A1 A8 A12 A14 A15	3	0	3	
	A21 B2 B3				
Personalized attention		1	0	1	
(*)The information in the planning table is for guidance only and does not take into account the betergeneity of the students					

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Methodologies				
Methodologies	Description			
Guest lecture /	Oral presentation, complemented by the use of audiovisual material and the interaction with the students, to introduce the			
keynote speech	basic contents of the subject to transmit knowledge and facilitate learning.			
Seminar	Activity to be developed in small groups.			
	In-depth study of the contents introduced in the keynote lectures.			
	Questions and problems related to the contents of the subject are discussed and/or are solved in group, with support and			
	direct supervision of the lecturer.			
	The activity to be carried out before and during each session are indicated prior to a face-to-face session.			
Laboratory practice	Computer practices developed at the informatic labs.			
	Practical problems related to the contents of Quantum Chemistry are solved by using computer software commonly used in			
	scientific calculations.			
	Students must solve and hand-in a questionnaire concerning the practical sessions.			
Supervised projects	Homework performed in groups aimed at helping students to work independently, under the guidance of the lecturer.			
	Activities related to the contents of the seminars are proposed. They must be solved in group and, subsequently, must be			
	explained to the instructor in a face-to-face session.			



Objective test	Written in-class exam to be held at mid-semester.				
	The test can combine multiple-choice questions and short answer questions.				
	The basic principles and the formulation of quantum theory and their applications to simple systems are evaluated.				
	The test is solved and discussed in a subsequent session.				
	It will serve as feed-back to both students and instructors to assess the progress of the teaching-learning process.				
Objective test	Written in-class exam to be held at the end of the semester.				
	The test can combine multiple-choice questions and short answer questions.				
	The basic principles and the application of quantum theory to atoms and molecules.				
	The test is solved and discussed in a subsequent session.				
	It will serve as feed-back to both students and instructors to assess the progress of the teaching-learning process.				
Mixed	Final written exam to be held at the end of semester, and in second chance in July.				
objective/subjective	Knowledge, understanding, reasoning and critical thinking are assessed.				
test	It will consist on a combination of different types of questions: multiple choice and/or short answer combined with problem				
	solving.				
	It will be held on the dates approved by the Faculty Board.				

Personalized attention				
Methodologies	Description			
Guest lecture /	Homework of supervised projects done by each group should be presented to the instructor in a tutoring session.			
keynote speech	In addition, students are encouraged to make use of the tutoring sessions to solve any doubt.			
Seminar	Tutoring schedule will be decided at lecturers and students convenience.			
Supervised projects	Part-time students and those with special academic leave permission could ask for presential or email tutorials when			
	necessary.			

Assessment			
Methodologies	Competencies	Description	Qualification
Objective test	A1 A8 A14	Written in-class exam at mid-semester.	7.5
		The basic principles and the formulation of quantum theory and their applications to	
		simple systems are assessed as short answer questions and/or multiple-choice	
		questions.	
Mixed	A1 A8 A12 A14 A15	Final written exam at the end of semester.	70
objective/subjective	A21 B2 B3	All contents of the subject are assessed.	
test		It has two different types of questions: short answer questions (50%) combined with	
		problem solving (50%).	
Laboratory practice	A1 A8 A21 C1 C3	Assessment of skills in solving problems of Quantum Chemistry by using computer	10
		software commonly used in scientific calculations.	
		Attendance to all scheduled computer lab. sessions is mandatory to pass the course.	
		The final written report is also assessed.	
Supervised projects	A1 A8 A16 B2 B3 B5	Assesment of teamwork skills in resolution of problems or questions related to the	5
	C1 C3	contents of the seminars.	
		Apart from the solution of the proposed activities, the active paticipation in the	
		face-to-face session also contribute to the assessment.	
Objective test	A1 A8 A14	Written in-class exam at the end of the semester.	7.5
		The application of quantum theory to atoms and molecues is assessed as short	
		answer questions and/or multiple-choice questions.	

Assessment comments



- It is compulsory the regular attendance to the keynotes and seminar sessions to attend computer practical sessions. - It is mandatory to obtain a minimum mark of 2.0 (out of a possible 5) in each of the parts of the final written exam and a minimum mark of 4.5 (out of a possible 10) to pass the course. - To pass the course, the final grade has to be equal to or greater than 5 (out of a possible 10) and the minimum score on the final written exam. If the average is equal to or greater than 5 (out of 10) but the threshold marks were not met, the final mark will be 4.5 (fail). - Students who has attended the practical sessions or the final exam will be assessed. Second opportunity of July- According to the rules contained in ?Probas de Avaliación e Actas de Cualificación de Grao e Mestrado?, the so-called ?second opportunity of July? is understood as a second opportunity to retake the final written exam. The mark of this second exam will be considered together with the others obtained during the course, corresponding to the other activities. The percentages of the different contributions will be the same as those of the former "first opportunity". -Mark Honors: priority is given in the first opportunity (January). Honors may only be granted in July if their number have not be exhausted in January final qualifications. Part-time students or students with special academic permission (according to the rules of the UDC): The same evaluation criteria listed above are applied, but it's not mandatory to attend classroom lectures, to hand-in the supervised projects and to attend the objective test. It is compulsory to attend computer practical sessions. It will be tried to fit the dates to the student's availability. The final grade will be the sum of 10% of the mark obtained in the practical sessions and 90% of the mark obtained in the mixed test. The same criteria will be applied to both opportunities. Students who has not attended the final exam will be assessed as "non attendance". Succesive academic years: - The teaching-learning process, including assessment, refers to an academic course and, therefore, will restart as new with every new academic year, including all activities and assessment procedures scheduled for that course. Mixed objective/subjective test dates: Final written exam for the two opportunities will be held on the

official dates approved by the Faculty Board.

Requirements to pass the course:

- Attendance to all scheduled computer lab sessions is mandatory to pass the course.

	Sources of information	
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Recommendations
Subjects that it is recommended to have taken before
Mathematics 1/610G01001
Mathematics 2/610G01002
Physics 1/610G01003
Physics 2/610G01004
General Chemistry 1/610G01007
General Chemistry 2/610G01008
Subjects that are recommended to be taken simultaneously
Subjects that continue the syllabus
Physical Chemistry 2/610G01017
Other comments

(*)The teaching guide is the document in which the URV publishes the information about all its courses. It is a public document and cannot be modified. Only in exceptional cases can it be revised by the competent agent or duly revised so that it is in line with current legislation.